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~~Molecular Geometry 3 9. Explain the difference between a bonding electron domain and a nonbonding  
electron domain using the examples in Model 1. 10. Circle the correct word or phrase to complete the  
sentences: Pairs of electrons will (attract/repel) each other.~~

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Molecular Geometry How can molecular shapes be predicted using the VSEPR theory? Why? When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The true shape of a molecule is important because it determines many physical and chemical properties for the substance

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The valence shell electron pair repulsion (VSEPR) theory (or "VESPER" for short) is how the geometry of a molecule is determined around a central atom. The molecular geometry main shapes are tetrahedral, trigonal planar, trigonal pyramidal, bent, and linear and are named by measuring the bond angles between the central atom and another atom...

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Molecular Geometry Vsepr Theory Worksheet Answers

13, Based on the information in the Read This! section, sketch one of the molecular shapes shown below in each of the boxes provided in Model 1. Trigonal planar 120° Bent 104.5° Three-Dimensional Molecular Shapes Tetrahedral 109.5° STOp Molecular Geometry Linear Pyramidal 107°

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